In spite of these difficulties, extrapolation to 1 atm gives a value of $K$ for sodium fluorenyl at ~25°C which is identical with that obtained by extrapolation to 25°C, of the previously reported data [1] pertaining to lower temperatures (i.e., $K_{25°,\ 1atm} \sim 0.058$). Some discrepancy is found, however, for the lithium salt (our $K = 2.7$, reported 4.6).

The equation for $K$ is modified if $\varepsilon_{loose}/\varepsilon_{tight} \neq 1$. Denoting this ratio by $\gamma$ we find

$$K = (R - r_1)/(\gamma - Rr_2).$$

It is possible that our $r_2$ is too small, because it was determined from the spectrum of lithium fluorenyl recorded at the highest pressure. This may account for our low value of $K$ for the lithium salt at 1 atm. At higher pressures the error is probably partly compensated by $\gamma$ being larger than 1.

Finally, we have checked that any reasonable changes of $r_1$ and $r_2$ have negligible effect upon the final values of $\Delta V$ for the sodium salt, although the results obtained for the lithium salt are less certain.

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